



Green reduction of graphene oxide using kaffir lime peel extract (Citrus hystrix) and its application as adsorbent for methylene blue by Ronald Wijaya

From Cek turnitin 1 (Pengecekan jurnal)

Processed on 21-Feb-2020 21:41 WIB

ID: 1261458969

Word Count: 6048

Similarity Index	Similarity by Source
12%	Internet Sources: 7% Publications: 12% Student Papers: N/A

sources:

- 1 2% match (publications)
["Handbook of Graphene", Wiley, 2019](#)

- 2 1% match (publications)
["A New Generation Material Graphene: Applications in Water Technology", Springer Science and Business Media LLC, 2019](#)

- 3 1% match (publications)
[Yuxi Xu, Lu Zhao, Hua Bai, Wenjing Hong, Chun Li, Gaoquan Shi. "Chemically Converted Graphene Induced Molecular Flattening of 5,10,15,20-Tetrakis\(1-methyl-4-pyridinio\)porphyrin and Its Application for Optical Detection of Cadmium\(II\) Ions", Journal of the American Chemical Society, 2009](#)

- 4 1% match (Internet from 05-May-2019)
<https://openchemicalengineeringjournal.com/VOLUME/10/PAGE/18/FULLTEXT/>

- 5 1% match (Internet from 31-Jan-2020)
http://ambi-agua.net/seer/index.php/ambi-agua/article/download/2159/pdf_1

- 6 1% match (publications)
[Asghar Faghihi, M.H. Vakili, Ghader Hosseinzadeh, Mousa Farhadian, Ziaeddin Jafari. "Synthesis and application of recyclable magnetic freeze-dried graphene oxide nanocomposite as a high capacity adsorbent for cationic dye adsorption", Desalination and Water Treatment, 2016](#)

- 7 1% match (Internet from 29-Dec-2017)
<http://scholarbank.nus.edu.sg/bitstream/10635/119833/1/FAN%20ZENG%20Final%20Thesis%20Graphene%20Aerog>

- 8 1% match (publications)
[Chengyang Li, Zechao Zhuang, Xiaoying Jin, Zuliang Chen. "A facile and green preparation of reduced graphene oxide using Eucalyptus leaf extract", Applied Surface Science, 2017](#)

- 9 < 1% match (Internet from 13-Nov-2018)
<https://zdoc.site/comparison-of-methylene-blue-adsorption-from-aqueous-solutio.html>

- 10 < 1% match (publications)
[Cheng, Zheng, Rendang Yang, and Xuehuan Zhu. "Adsorption behaviors of the methylene blue dye onto modified sepiolite from its aqueous solutions", Desalination and Water Treatment, 2016.](#)

-
- 11** < 1% match (publications)
[Jana, Milan, Sanjit Saha, Partha Khanra, Naresh Chandra Murmu, Suneel Kumar Srivastava, Tapas Kuila, and Joong Hee Lee. "Bio-reduction of graphene oxide using drained water from soaked mung beans \(*Phaseolus aureus* L.\) and its application as energy storage electrode material", *Materials Science and Engineering B*, 2014.](#)
-
- 12** < 1% match (Internet from 09-Feb-2020)
<https://rd.springer.com/article/10.1007/s10904-013-9927-5>
-
- 13** < 1% match (publications)
[Abu Nasar, Fouzia Mashkoor. "Application of polyaniline-based adsorbents for dye removal from water and wastewater—a review", *Environmental Science and Pollution Research*, 2019](#)
-
- 14** < 1% match (publications)
[Essam Hares, A. H. El-Shazly, M. F. El-Kady, A. S. Hammad. "Electrophoretic Deposition of Graphene Oxide Nanosheets on Copper Pipe for Corrosion Protection", *Arabian Journal for Science and Engineering*, 2019](#)
-
- 15** < 1% match (publications)
[Kirti Thakur, Balasubramanian Kandasubramanian. "Graphene and Graphene Oxide-Based Composites for Removal of Organic Pollutants: A Review", *Journal of Chemical & Engineering Data*, 2019](#)
-
- 16** < 1% match (publications)
[Wang, Ya, Lin Zhu, Haitao Jiang, Fang Hu, and Xiangqian Shen. "Application of longan shell as non-conventional low-cost adsorbent for the removal of cationic dye from aqueous solution", *Spectrochimica Acta Part A Molecular and Biomolecular Spectroscopy*, 2016.](#)
-
- 17** < 1% match (publications)
[Yurong Liu, Rong Hu, Zongqiang Zhang. "A facile colloidal crystal templating method to produce three-dimensional hierarchical porous graphene–Fe₃O₄ nanocomposite for the removal of dyes from aqueous solution", *Journal of Porous Materials*, 2018](#)
-
- 18** < 1% match (Internet from 26-Jan-2020)
<https://www.mdpi.com/2071-1050/10/6/1795/htm>
-
- 19** < 1% match (Internet from 16-Feb-2020)
https://www.medscape.com/viewarticle/889333_1
-
- 20** < 1% match (Internet from 26-May-2019)
<http://www.icontrolpollution.com/articles/preparation-and-characterisation-of-activated-carbon-from-delonix-regia-seeds-for-the-removal-of-methylene-blue-dye-.php?aid=82942&view=mobile>
-
- 21** < 1% match (Internet from 19-Oct-2019)
<https://www.tandfonline.com/doi/pdf/10.1080/19443994.2012.692037>
-
- 22** < 1% match (Internet from 01-Feb-2020)
<https://www.degruyter.com/view/j/polyeng.2019.39.issue-5/polyeng-2018-0275/polyeng-2018-0275.xml>

23 < 1% match (publications)

[Wenny Irawaty, Lannie Hadisoewignyo, Jenifer Wijaya, Shindy Saera Sababalat, Andy Sebastian Tanjung. "Design of kaffir lime peel extract tablets by direct compression: Effects of filler-binder on rheological properties", AIP Publishing, 2019](#)

paper text:

OPEN Green Reduction of Graphene Oxide using Kaffir Lime Peel Extract (*Citrus hystrix*) and Its Application as Adsorbent for Methylene Blue Ronald Wijaya 1, Gracia Andersan1, Shella Permatasari Santoso1,2 & Wenny Irawaty1* Green

1reduction of graphene oxide (GO) by phytochemicals was

explored using the aqueous extract of kaffir lime peels. The research methods included preparation of extracts, preparation of GO, preparation and characterization of reduced-GO (RGO) using Fourier Transform Infrared (FTIR), X-ray diffraction (XRD), and UV-Vis spectroscopy, as well as methylene blue (MB) adsorption test using RGO. The RGO characterization showed that GO was successfully reduced by a C=C group restoration. The MB adsorption kinetics profile in RGO is more suitable for the pseudo-second-order model, whereas for the adsorption isotherm it is more suitable for the Langmuir model with a maximum adsorption capacity (q_{max}) of 276.06 mg/g at room temperature. The best ratio of GO: kaffir lime peel extract used to prepare RGO was at a ratio of 1: 2.

6Based on the ΔG , ΔH , and ΔS values, the adsorption of

RGO-MB was defined as spontaneous and endothermic process. The results promise the potential application of RGO derived via green route to remove cationic dye in wastewater. Methylene Blue (MB) is a cationic dye that has a dark green color¹. MB is very soluble in water and alcohol and often used as the coloring agent for textiles². Methylene blue can have a negative impact on human body³. If inhaled, it can cause respiratory disease, when MB is accidentally swallowed, it will produce a burning sensation and can cause nausea, vomiting, diarrhea, gastritis, abdominal and chest pain, severe headaches, lots of sweating, mental confusion and methemoglobinemia⁴. Therefore, the presence of MB in the water needs to be well addressed. There are several methods used to treat waste from textile water pollution, by ultrafiltration method⁵, ozonization⁶, ion exchange⁷, photocatalysis⁸, and adsorption⁹. Of the many methods, adsorption the most promising method for dealing with textile wastewater; using commercial adsorbents such as activated carbon¹⁰, zeolites¹¹ and nanomagnetic materials^{11–13}. In this study, the concentration of MB in simulated textile wastewater was reduced by the adsorption

1onto the surface of graphene oxide- based material. Graphene Oxide

(GO) is a carbon nanomaterial that can make strong interaction with organic molecules through hydrogen bonds and π - π interaction¹⁴. GO has a stable 2-dimensional structure as electrical and thermal conductors¹⁵. GO is widely used in many applications such as sensor¹⁶, catalyst¹⁷, electronics¹⁸, water splitting¹⁹, CO₂ reduction²⁰, and water treatment^{17,21}. The ability of GO to adsorb MB is tremendous, with removal efficiency up to 98.8%. Some modifications can be applied to enhance the ability of GO in MB adsorption, such as composite photocatalyst^{17,22,23}. A further modification to improve the performance of GO-based materials is essential. GO is heavily decorated by the oxygen functional group (C-O), and therefore if the oxygen groups can be reduced, the surface area may be expanded. In this state, GO is turned into Reduced Graphene Oxide (RGO). The conventional method to reduce GO is performed using

chemicals such as hydrazine²⁴, sodium borohydride²⁵, hydroxylamine²⁶, etc. The chemical synthesis route offers an efficient method, but the chemicals used are not environmentally friendly. The decrease of RGO performance due to the employment of strong reducing agent during the process has been reported²⁷. Thermal treatment is another method to reduce GO^{28,29}. The heat treatment can efficiently

1 remove the oxygen- containing **functional groups** and thus, open **the**

possibility to fabricate RGO. However, the system

1 requires a high amount of energy

as well as complicated experimental set-up. ¹Department

4 of Chemical Engineering, Widya Mandala Surabaya Catholic University, Kalijudan 37, Surabaya, 60114, Indonesia. ²Department of Chemical Engineering, National Taiwan University of Science and Technology, No. 43, Sec. 4, Keelung Rd, Da'an District, Taipei City, Taiwan.

*email: wenny_i_s@ukwms.ac.id Electrochemical reduction offers faster and safer route compared to the previous methods, but its reduction efficiency is reported less than the hydrazine reduction³⁰. Among those methods, the reduction using chemical agents is still the most effective technique for producing RGO. However, the use of hazardous chemicals during RGO can be a problem for the environment²⁷. Some natural compounds have been introduced to replace the use of the hazardous reducing agent, i.e. extract of eucalyptus leaves³¹, ascorbic acid³², green tea³³, lemon³⁴, etc. However, most bioreducing agents used in RGO preparation are food sources; this raises the possibility of food competition. Therefore, there is still a need to explore non-food sources as reducing agents; agricultural by-products such as peels, seeds, etc, can be a potential source. Herein, we present the RGO preparation using a reducing agent prepared from kaffir lime peel. The reduction of GO using kaffir lime peel extract has not been reported in any literature. Our approach offers several advantages such as eco-friendly reducing agent, simple extract preparation, and non-toxic wastes generated at the end of reduction process. **Materials and Methods** Natural graphite powder was supplied by PT. Brataco (Surabaya, Indonesia). Hydrochloric acid (SAP, 37%), sulfuric acid (H₂SO₄, CAS 7664-93-9, 99.999%), hydrogen peroxide (H₂O₂, SAP, 30%), potassium permanganate (KMnO₄, SAP, ≥99.0%), and methylene blue (C₁₆H₁₈ClN₃S, CAS 122965-43-9) were used as received.

23 Kaffir lime fruits were obtained **from local** kaffir lime plantation **in Lumajang (East Java,**

Indonesia). Preparation

7 of graphene oxide (GO). GO was prepared by using the **modified**

Hummer's method¹⁸. The preparation procedure was briefly given as follows: 3 g of graphite powder was put into 70 mL of H₂SO₄ solution under ice bath conditions with the temperature was set below 20 °C, then 9 g of KMnO₄ was added to the mixture. Subsequently, the mixture was transferred to an oil bath (Mimmert), and the temperature was maintained at 40

1°C. During the reaction process, the mixture was continuously stirred at 600 rpm for 30 min. When the

reaction was completed,

3150 ml of RO water was added, and the mixture constantly stirred at 600 rpm for 15 min within 95 °C. Then, 500 mL of water was added, followed by slow addition of 15 mL of H₂O₂ (30%) to remove the residue. A change of solution color was observed from brown to yellowish. The suspension is then filtered and washed with

250 mL of dilute HCl solution (1:10). The resulting solid was dried naturally, followed by adding 600 mL of water. The GO solution was then

11 repeatedly washed with water to remove all residual salts and acids.

The neutralized GO solution was stirred overnight, followed by the sonication process to exfoliate the graphite oxide into graphene oxide. The GO dispersion was then stored for the reduction process. Preparation of kaffir lime peel extract (*Citrus hystrix*). Twenty-five grams of kaffir lime peels were extracted with 1000 mL ethanol (41%)

1 at room temperature for 8 h. The extract was then

filtered to separate the solid part and kept in closed container to be further used for GO reduction. Total Phenolic Content (TPC) test was carried out to determine the amount of phenolics in the extract of kaffir lime peels. Reduction of GO by kaffir lime peel extract (*Cytrus hystrix*). GO and kaffir lime peel extract was mixed with the GO/extract ratio (v/v) 1:1, 1:2, and 1:4 under vigorous stirring (600 rpm) for 8 h. Subsequently, the mixtures were washed and sonicated for 30 min. This treatment was repeated

8 several times until a clear solution was obtained. Then, the materials were dried in a vacuum

oven (Vacucell, 55). Characterization. Ultraviolet

11-visible (UV-vis) spectroscopy of GO and RGO was recorded on a

Shimadzu UVMINI-1240 spectrophotometer. Crystalline properties of graphite, GO, and RGO were analyzed using X-Ray Diffraction (XRD, Philips X'pert Xray Diffractometer) operated at 40 kV; 30 mA using Cu-K α 1 (λ = 0.15406) at 0.02° step size and 10.16 s step time from 5° to 60°. The surface functional group analysis of the graphite, GO, and RGO was analyzed using Fourier Transform Infrared (FTIR, Shimadzu 8400 S) using KBr powder method. The FTIR spectra were obtained at wavenumber from 1000 to 4000 cm⁻¹. Preparation of methylene blue (MB) solutions. A stock solution of 1000 mg/L MB was made from MB powder, which was dissolved in water. The solution used in the experiment was prepared by diluting the stock solution from MB

to 50 mg/L. The pH of the solution measured using pH meter (Mettler Toledo). The experiment of MB adsorption. The contact time effect on MB adsorption was carried out by adding 5 mg RGO

10 into 20 mL of MB solution with a concentration of 50 mg/L

in a 100 mL erlenmeyer flask. The flasks were placed in a water bath shaker at 30 °C (Memmert SV-1422) and shaken until the equilibrium time was reached. At the specified time, the RGO and MB

9 were separated by centrifugation at 4500 rpm for 10 min. The

adsorption isotherm experiment was conducted at room temperature using three variations of RGO (1:1, 1:2, and 1:4). The experiment was carried by different addition mass of adsorbent (2–12 mg, step size: 2 mg) into a series of erlenmeyer flasks, where each flask contains 20 mL of MB solution with a concentration of 50 mg/L. These flasks were shaken at 90 rpm until the equilibrium state was achieved. The RGO and MB

9 were separated by centrifugation at 4500 rpm for 10 min. The

final

20 concentration of MB was analyzed by UV-Vis spectrophotometer at 665 nm.

The adsorption capacity of RGO was calculated by: $q_t = \frac{m}{V} (C_0 - C_t)$

– C_t (mg/L) is the concentration of MB at time t , m (g) is the adsorbent quantity, and V (L) is the volume of MB solution. (1) Figure 1. UV-Vis Spectrum of GO and RGO. Figure 2. FT-IR spectra of graphite, GO, and RGO. where C_0

16 (mg/L) is an initial concentration of MB, C_t (mg/L) is the concentration of MB at time t ,

m (g) adsorbent quantity, and V (L) is the volume of MB solution. At a constant value of C_t , the parameter of C_t and q_t becomes C_e and q_e , where C_e (mg/L) is the equilibrium concentration of MB. Results and Discussion Characterization of graphene materials. UV-Vis spectroscopic analysis was carried out to monitor changes in functional groups from GO to RGO. It can be seen from Fig. 1 that the peak transition $\pi \rightarrow \pi^*$ for graphitic groups on GO was detected at a wavelength of ~300 nm. The shoulder peak at ~343 nm representing $n \rightarrow \pi^*$ for the C=O group, indicates the presence of a C-O functional group in GO³². The GO transformation into RGO was confirmed by loss of the shoulder peak, where the C-O group in RGO has been reduced. The transformation of GO to RGO was also supported by the change of GO solution color, which was initially brown to blackish after being reduced by kaffir lime peels extract. The surface functional groups of graphite,

1 GO, and RGO were obtained by Fourier Transform Infrared (FTIR) analysis. Fig. 2

shows

1the presence of several functional groups in

the three samples tested. The results of FTIR analysis summarized in Table 1 showed the presence of C=C groups in graphite at the peak of 1512.08 cm^{-1} . The formation of GO was indicated by the presence of a hydroxyl group at 3656.54

14 cm^{-1} , a carbonyl group (C=O) at 1704.73 cm^{-1} ,

an alkene group (=CH) at 1090.23

14 cm^{-1} , and ester group (C-O) at 1382 cm^{-1} . The GO

reduction process into RGO caused the aromatic group (C=C) restored by the emergence of 1456 cm^{-1} and 1735.22 cm^{-1} peaks in RGO. The change of the peak indicates that the GO was successfully reduced to RGO³⁵. FTIR analysis shows that there are still C=O groups remained on the reduced GO; this signifies that the GO reduction partially occurred. However, it is still unclear as to whether C=O persists even after suffering a reduction. A similar phenomenon was experienced by Aliyev et al., where the RGO produced still has C=O groups remaining even after reduction using hydrazine monohydrate³⁶. Dreyer and group suggest that in the process of GO reduction, the reducing agent can breakdown the epoxy groups of GO and produce carbonyl groups³⁷. The

7X-Ray Diffraction (XRD) analysis was conducted to determine the crystallinity of graphite, GO, and RGO. The peak point of the GO

is determined based on the degree of oxidation and the exfoliation process. It can be Functional Group Alkane C-C Wavenumber ($1/\text{cm}$) Graphite 1502.73 GO — RGO — Hydroxyl (O-H) — 3656.54 3656.54 Carbonyl (C=O) — 1704.73 1706.85 Alkene (=C-H) — 1090.23 1087.57 Ester (C-O) — 1382 — Aromatic (C=C) — — 1456 Aromatic (C=C) — — 1735.22 Table 1. The surface functional groups of Graphite, GO and RGO. Sample 2θ (degree) Spacing-d (\AA) Graphite 26.45 3.37 GO 10.11 8.75 RGO 8.75 10.10 26.34 3.38 Table 2. Peak result of XRD pattern for

8Graphite, GO, and RGO. Figure 3. XRD patterns of (a) graphite (b) GO and RGO. seen from Fig.

3 and Table 2 that graphite, which initially had a sharp graphitic peak at 26.45° and spacing of 3.35 \AA . After being synthesized became GO, the graphitic peak in 26.45° was found disappear, and new peak appeared at 10.11° with lots of diffraction. This indicates that synthesized graphite into GO undergoes peak shifting due to the oxidation and intercalation process of epoxy, hydroxyl, etc.³². The spacing or distance between layers that were initially 3.37 \AA changed to 8.75 \AA also indicates an increasing degree of oxidation, and the GO exfoliation process is getting stronger³⁸. After being reduced being RGO it was found that the peak point in the area of 10.11° has broadened, and the graphitic peak intensity at 26.34° is increasing. Thus it can be indicated that the reduction process, there is a degradation in C-O groups and the restoration of the C=C bond structure. Adsorption kinetic study. The adsorption kinetic curve of MB over RGO

17 are shown in Fig. 4. It can be seen that the adsorption

equilibrium time was reached after 420 min.

15 The adsorption kinetics parameter was determined using the pseudo-first and the pseudo-second-order models.

These models

17 are calculated using the following equations:

pseudo - first $\log(q_e - q_t) = \log q_e - k_1 t$ (2) pseudo - second $t = 1 + t q_t k_2 q_e^2 q_e$ (3) The parameters of the adsorption kinetics of MB over RGO are summarized in Table 3. Based on the sum square error (SSE), physisorption is the dominant process controlled the adsorption of MB on RGO. In reverse, if the model obtained refers to pseudo-second-order, then RGO-MB will tend to the mechanism of chemisorption³⁹. The equilibrium state of MB adsorption on RGO obtained in kinetic experiment is used as reference time in adsorption isotherm process. Figure 4. Adsorption kinetics of MB adsorption onto RGO. Kinetic Model Parameters q_e (mg/g) Value 192.79 Pseudo-first-order k_1 (1/min) SSE 0.0032 606.53 Pseudo-second-order q_e (mg/g) k_2 (g/mg min) $\times 10^{-6}$ SSE 288.11 8.12 582.92

22 Table 3. Parameters of the pseudo-first-order and pseudo-second-order kinetic model for MB adsorption.

18 From the pseudo-first-order model, the q_e

value is 192.79 mg/g, and the k_1 value is 0.0032 min^{-1} with SSE of 606.53. Meanwhile, the pseudo-second-order equation obtained q_e of 288.11 mg/g and k_2 value of 8.12 $((\text{g}/\text{mg min}) \times 10^{-6})$ with SSE of 582.92. From the comparison of SSE values in the kinetic model used, the RGO adsorption kinetics is more in line with the pseudo-second-order, which means that the adsorption process of RGO-MB was more controlled by chemisorption. Isotherm adsorption study. The adsorption isotherm of RGO was carried out to study the interaction between adsorbate and adsorbent.

13 Adsorption isotherm was carried out with

a span of 420 min, where this time is equilibrium time as obtained from the adsorption kinetics data. The adsorption isotherm was modeled using the Langmuir and Freundlich models. The Freundlich and Langmuir model can be represented mathematically as follows: Freundlich $q_e = K_f C_e^{1/n}$ (4) Langmuir $q_e = q_{\text{max}} \frac{K_L C_e}{1 + K_L C_e}$ (5) In the Freundlich model, K_f represents the adsorption capacity of the adsorbent, and n showed the heterogeneity and adsorption intensity of RGO. In the Langmuir model, q_{max}

10 represents the maximum adsorption capacity, K_L states the Langmuir constant. The result of isotherm adsorption

of RGO-MB is presented in Fig. 5a and Fig. 5c, and the calculated parameters are summarized in Table 4. Based on the SSE obtained, it can be seen the SSE of the Langmuir model is smaller than the SSE of the Freundlich model. Also, a small n value indicates that the adsorbent heterogeneity is minimal and tends to be homogeneous. This shows the uniformity of the adsorption surface that supports conformity with the Langmuir model. The affinity of the adsorption between RGO-MB is determined by the KL parameter where the value is very small; which supports the results of data n from the Freundlich model, where small n is indicating that the adsorbent more less intensity to adsorb the adsorbate. It can be concluded that the Langmuir isotherm model is the most suitable model for the MB adsorption process using RGO. The process of adsorption of MB by RGO studied tends to be slow. Table 4 shows the relationship of the variation in concentration to the adsorption capacity. At a ratio of 1:1, the adsorption capacity (q_{max}) was found 228.07 mg/g, at a concentration of 1:2 q_{max} , the higher adsorption capacity was obtained (276.06 mg/g) and decreased at a ratio of 1:4. This can be possible that the CO group present in the GO has not been completely reduced in the concentration of 1:1, whereas at the ratio of 1:4 the reduced GO is at Figure 5. Adsorption isotherm of MB adsorption onto RGO. (a) Freundlich model at different ratio of GO/ extract (b) Freundlich model at different temperatures (c) Langmuir model at different ratio of GO/extract (d) Langmuir model at different temperatures. Adsorbent RGO 1:1 Freundlich n 1.57 KF (L/g) 8.23 SSE 111.73 Langmuir q_{max} (mg/g) KL (L/g) SSE 228.07 0.015 104.41 RGO 1:2 1.49 7.98 227.54 276.06 0.013 213.06 RGO 1:4 1.57 7.29 51.44 209.34 0.014 39.00 Table 4. Parameters of Freundlich and Langmuir isotherm adsorption model for MB adsorption. the saturated condition. The adsorption capacity of MB using various RGO suggests that RGO 1:2 has the best performance to remove MB. The ability of RGO to adsorb MB was compared to other studied adsorbents, as shown in Table 5. RGO produced in this study exhibited relative higher of adsorption capacity compared to other adsorbents. For the same target adsorbate (MB), up to 276 mg of MB can be adsorbed on the surface of RGO; which

18 is much higher than the adsorption capacity of

nanocomposite of β -cyclodextrin/magnetic graphene oxide, that is only 94 mg/g. The photocatalyst Fe₃O₄-RGO-TiO₂, only able to adsorb 1.67 mg MB per gram of particle. However, the use of activated seagrass Posidonia oceanica waste can achieve almost ten times higher than our adsorbent⁴⁰. The high adsorption capacity of RGO (this study) can be explained by its high specific surface area, i.e. 3127 m²/g as measured

1 by the Brunauer-Emmett-Teller method. The significantly increased of the

specific surface area from GO (832 m²/g) to RGO has facilitated the MB adsorption. This result suggests that kaffir lime peel extract is potent to be used as the bioreducing agent to modify GO for MB adsorption. Further study to elucidate the mechanism of reduction activity performed by phytochemicals of kaffir lime peels is required. Adsorption thermodynamics. Adsorption thermodynamics proceeds to evaluate the effect of temperature due to adsorption process of RGO-MB. The thermodynamics parameter is calculated using following equations: $\Delta G = -RT \ln KL$ $\Delta G = \Delta H = \Delta S$. $T \ln(KL) = \ln \left(\frac{Q_{Cee}}{C_{e0}} \right) = \frac{\Delta S}{R} - \frac{\Delta H}{RT}$ (6) (7) (8) 6 Adsorbents Graphene/SrAl₂O₃:Bi³⁺ Adsorption capacity (mg/g) 42.92 References 17 β -cyclodextrin/MGO 93.97 43 g-C₃N₄(Melamine) 1.64 44 g-C₃N₄(Thiourea) 1.87 44 g-C₃N₄(Urea) 2.51 44 TiO₂/Na-g-C₃N₄ 1.8 45 Magnetic carboxyl functional nanoporous polymer 57.74 46 Activated Posidonia oceanica waste 2,681.90 40 CeO₂ 4.37 47 Fe₃O₄-RGO-TiO₂ 1.67 48 Ag-Fe₃O₄-polydopamine 45 23 RGO 1:2 276.06 Present work Table 5. Adsorption capacity of several adsorbents to adsorb methylene blue. T (K) 303

5 ΔG (kJ mol⁻¹) ΔH (kJ mol⁻¹) ΔS (J mol⁻¹ K⁻¹) -5.98 19 .15

82.93 313 –6.81 333 –8.47 Table 6. Thermodynamics parameters for MB adsorption onto RGO. Figure 6. MB adsorption mechanism using RGO.

12where R is the universal gas constant ($8.314 \text{ J mol}^{-1} \text{ K}^{-1}$) and T is absolute temperature (K). ΔH (kJ mol^{-1}), ΔS

($\text{J mol}^{-1} \text{ K}^{-1}$), and ΔG (kJ mol^{-1}) are enthalpy, entropy, and Gibbs free energy, respectively.

6 ΔH and ΔS were calculated from the slope and intercept of $\ln(KL)$ versus $1/T$ which was represented in Fig. 5b, and

Fig. 5d. As seen in Table 6, the ΔG value was found to be $-5.98 \text{ kJ mol}^{-1}$ at 303 K and decrease to $-8.47 \text{ kJ mol}^{-1}$ at 333 K. The negative values of ΔG indicates the MB adsorption onto RGO is spontaneous. The value of ΔH and ΔS was found to be 19.15

21 kJ mol^{-1} and $82.93 \text{ J mol}^{-1} \text{ K}^{-1}$, respectively. The positive value of ΔH

and ΔS can be related to the adsorption mechanism of RGO-MB41 as seen in Fig. 6, where the positive value of ΔH also represents that RGO-MB adsorption process was endothermic. The high adsorption capacity of RGO is attributed to π - π interaction between MB molecule and the surface of RGO14. The high area exhibited by RGO favors the π - π electron donor interaction between the molecule of MB and RGO. Accordingly, the adsorption of MB on the surface of RGO could go up to 276 mg/g . In addition, electrostatic forces caused by difference in surface charge between MB and RGO also support the adsorption process. In aqueous solution, MB will exist as positively charged ions, while RGO tends to be negatively charged in water and thus, this has facilitated the electrostatic force between the two⁴². Conclusions The reduction: of GO to RGO by using kaffir lime peel extract as a bioreductor was successfully carried out. The ratio of bioreductor and reduction time has been observed to influence the characteristics of RGO, and the optimal conditions were found to be 1:2 and 8 h, respectively. Analyses of UV-Vis Spectroscopy, FTIR, and XRD indicated the reduction of GO to RGO by the C=C group restoration. The adsorption study performed on MB showed that the maximum adsorption capacity (q_{max}) of RGO was found 276.06 mg/g . MB adsorption by RGO followed pseudo-second-order model; whereas, for the isotherm, the system prefers in line with the Langmuir model. The RGO-MB adsorption interaction is more dominantly controlled by chemisorption. The adsorption of RGO-MB

15was a spontaneous and endothermic process. The

present study indicates RGO prepared by bioreduction of RG using kaffir lime peels extract a potential adsorbent to remove MB and therefore, it can be further developed for industrial applications. Received: 3 October 2019; Accepted: 31 December 2019; Published: xx xx xxxx References 1. Shahwan, T. et al. Green synthesis of iron nanoparticles and their application as a Fenton-like catalyst for the degradation of aqueous cationic and anionic dyes. Chem. Eng. J. 172, 258–266, <https://doi.org/10.1016/j.cej.2011.05.103> (2011). 2. Tiwari, J. N. et al. Reduced graphene oxide-based hydrogels for the efficient capture of dye pollutants from aqueous solutions. Carbon 56, 173–182, <https://doi.org/10.1016/j.carbon.2013.01.001> (2013). 3. El Qada, E. N., Allen, S. J. & Walker, G. M. Adsorption of methylene blue onto activated carbon produced from steam activated bituminous coal: A study of equilibrium adsorption isotherm. Chem. Eng. J. 124, 103–110, <https://doi.org/10.1016/j.cej.2006.08.015> (2006). 4. Ghosh, D. & Bhattacharyya, K. G. Adsorption of

methylene blue on kaolinite. *Appl. Clay Sci.* 20, 295–300, [https://doi.org/10.1016/S0169-1317\(01\)00081-3](https://doi.org/10.1016/S0169-1317(01)00081-3) (2002). 5. Vinodhini, P. A. & Sudha, P. N. Removal of heavy metal chromium from tannery effluent using ultrafiltration membrane. *Textiles and Clothing Sustainability* 2, 5, <https://doi.org/10.1186/s40689-016-0016-3> (2016). 6. Yurteri, C. & Gurol, M. D. Removal of dissolved organic contaminants by ozonation. *Environ. Prog.* 6, 240–245, <https://doi.org/10.1002/ep.670060411> (1987). 7. Padmavathi, R., Minnoli, M. & Sangeetha, D. Removal of heavy metal ions from waste water using anion exchange polymer membranes. *Int. J. Plast. Technol.* 18, 88–99, <https://doi.org/10.1007/s12588-014-9067-y> (2014). 8. Wang, C.-C., Li, J.-R., Lv, X.-L., Zhang, Y.-Q. & Guo, G. Photocatalytic organic pollutants degradation in metal-organic frameworks. *Energ. Environ. Sci.* 7, 2831–2867, <https://doi.org/10.1039/C4EE01299B> (2014). 9. Lee, J.-Y., Chen, C.-H., Cheng, S. & Li, H.-Y. Adsorption of Pb(II) and Cu(II) metal ions on functionalized large-pore mesoporous silica. *Int. J. Environ. Sci. Technol.* 13, 65–76, <https://doi.org/10.1007/s13762-015-0841-y> (2016). 10. Singh, K. P., Mohan, D., Sinha, S., Tondon, G. S. & Gosh, D. Color removal from wastewater using low-cost activated carbon derived from agricultural waste material. *Ind. Eng. Chem. Res.* 42, 1965–1976, <https://doi.org/10.1021/ie020800d> (2003). 11. Meshko, V., Markovska, L., Mincheva, M. & Rodrigues, A. E. Adsorption of basic dyes on granular activated carbon and natural zeolite. *Water Res.* 35, 3357–3366, [https://doi.org/10.1016/S0043-1354\(01\)00056-2](https://doi.org/10.1016/S0043-1354(01)00056-2) (2001). 12. Afkhami, A., Saber-Tehrani, M. & Bagheri, H. Modified maghemite nanoparticles as an efficient adsorbent for removing some cationic dyes from aqueous solution. *Desalin.* 263, 240–248, <https://doi.org/10.1016/j.desal.2010.06.065> (2010). 13. Qadri, S., Ganoe, A. & Haik, Y. Removal and recovery of acridine orange from solutions by use of magnetic nanoparticles. *J. Hazard. Mater.* 169, 318–323, <https://doi.org/10.1016/j.jhazmat.2009.03.103> (2009). 14. Li, Y. et al. Comparative study of methylene blue dye adsorption onto activated carbon, graphene oxide, and carbon nanotubes. *Chem. Eng. Res. Des.* 91, 361–368, <https://doi.org/10.1016/j.cherd.2012.07.007> (2013). 15. Pei, S. & Cheng, H.-M. The reduction of graphene oxide. *Carbon* 50, 3210–3228, <https://doi.org/10.1016/j.carbon.2011.11.010> (2012). 16. Cai, L. et al. Synthesis of Fe₃O₄/graphene oxide/pristine graphene ternary composite and fabrication electrochemical sensor to detect dopamine and hydrogen peroxide. *Chem. Phys. Lett.* 736, 136797, <https://doi.org/10.1016/j.cplett.2019.136797> (2019). 17. Oliva, J. et al. Flexible graphene composites for removal of methylene blue dye-contaminant from water. *Appl. Surf. Sci.* 436, 739–746, <https://doi.org/10.1016/j.apsusc.2017.12.084> (2018). 18. Chen, J., Yao, B., Li, C. & Shi, G. An improved Hummers method for eco-friendly synthesis of graphene oxide. *Carbon* 64, 225–229, <https://doi.org/10.1016/j.carbon.2013.07.055> (2013). 19. Kuang, P., He, M., Zou, H., Yu, J. & Fan, K. 0D/3D MoS₂-NiS₂/N-doped graphene foam composite for efficient overall water splitting. *Appl. Catal. B.* 254, 15–25, <https://doi.org/10.1016/j.apcatb.2019.04.072> (2019). 20. Bie, C., Zhu, B., Xu, F., Zhang, L. & Yu, J. In situ grown monolayer N-doped graphene on CdS hollow spheres with seamless contact for photocatalytic CO₂ reduction. *Adv. Mater.* 31, 1902868, <https://doi.org/10.1002/adma.201902868> (2019). 21. Zheng, Y., Cheng, B., You, W., Yu, J. & Ho, W. 3D hierarchical graphene oxide-NiFe LDH composite with enhanced adsorption affinity to Congo red, methyl orange and Cr(VI) ions. *J. Hazard. Mater.* 369, 214–225, <https://doi.org/10.1016/j.jhazmat.2019.02.013> (2019). 22. Yang, X. et al. Rapid degradation of methylene blue in a novel heterogeneous Fe₃O₄@rGO@TiO₂-catalyzed photo-Fenton system. *Sci. Rep.* 5, 10632, <https://doi.org/10.1038/srep10632> (2015). 23. Wu, M., Li, Y., Yue, R., Zhang, X. & Huang, Y. Removal of silver nanoparticles by mussel-inspired Fe₃O₄@ polydopamine core-shell microspheres and its use as efficient catalyst for methylene blue reduction. *Sci. Rep.* 7, 42773, <https://doi.org/10.1038/srep42773> (2017). 24. Park, S. et al. Hydrazine-reduction of graphite- and graphene oxide. *Carbon* 49, 3019–3023, <https://doi.org/10.1016/j.carbon.2011.02.071> (2011). 25. Shin, H.-J. et al. Efficient reduction of graphite oxide by sodium borohydride and its effect on electrical conductance. *Adv. Funct. Mater.* 19, 1987–1992, <https://doi.org/10.1002/adfm.200900167> (2009). 26. Zhou, X. et al. Reducing graphene oxide via hydroxylamine: A simple and efficient route to graphene. *J. Phys. Chem. C* 115, 11957–11961, <https://doi.org/10.1021/jp202575j> (2011). 27. Kuila, T. et al. Chemical functionalization of graphene and its applications. *Prog. Mater. Sci.* 57, 1061–1105, <https://doi.org/10.1016/j.pmatsci.2012.03.002> (2012). 28. Zhu, Y. et al. Exfoliation of graphite oxide in propylene carbonate and thermal reduction of the resulting graphene oxide platelets. *ACS Nano* 4, 1227–1233, <https://doi.org/10.1021/nn901689k> (2010). 29. McAllister, M. J. et al. Single sheet functionalized graphene by oxidation and thermal expansion of graphite.

Chem. Mater. 19, 4396–4404, <https://doi.org/10.1021/cm0630800> (2007). 30. An, S. J. et al. Thin film fabrication and simultaneous anodic reduction of deposited graphene oxide platelets by electrophoretic deposition. *J. Phys. Chem. Lett.* 1, 1259–1263, <https://doi.org/10.1021/jz100080c> (2010). 31. Jin, X., Li, N., Weng, X., Li, C. & Chen, Z. Green reduction of graphene oxide using eucalyptus leaf extract and its application to remove dye. *Chemosphere* 208, 417–424, <https://doi.org/10.1016/j.chemosphere.2018.05.199> (2018). 32. De Silva, K. K. H., Huang, H.-H. & Yoshimura, M. Progress of reduction of graphene oxide by ascorbic acid. *Appl. Surf. Sci.* 447, 338–346, <https://doi.org/10.1016/j.apsusc.2018.03.243> (2018). 33. Tai, M. J.-Y. et al. Green synthesis of reduced graphene oxide using green tea extract. *AIP Conference Proceedings* 2045, 020032, <https://doi.org/10.1063/1.5080845> (2018). 34. Hou, D., Liu, Q., Cheng, H. & Li, K. Graphene synthesis via chemical reduction of graphene oxide using lemon extract. *J. Nanosci. Nanotechnol.* 17, 6518–6523, <https://doi.org/10.1166/jnn.2017.14426> (2017). 35. Jana, M. et al. Bio-reduction of graphene oxide using drained water from soaked mung beans (*Phaseolus aureus* L.) and its application as energy storage electrode material. *Mater. Sci. Eng. B.* 186, 33–40, <https://doi.org/10.1016/j.mseb.2014.03.004> (2014). 36. Aliyev, E. et al. Structural characterization of graphene oxide: Surface functional groups and fractionated oxidative debris. *Nanomaterials* 9, 1180, <https://doi.org/10.3390/nano9081180> (2019). 37. Dreyer, D. R., Park, S., Bielawski, C. W. & Ruoff, R. S. The chemistry of graphene oxide. *Chem. Soc. Rev.* 39, 228–240, <https://doi.org/10.1039/b917103g> (2010). 38. Abdolhosseinzadeh, S., Asgharzadeh, H. & Seop Kim, H. Fast and fully-scalable synthesis of reduced graphene oxide. *Sci. Rep.* 5, 10160, <https://doi.org/10.1038/srep10160> (2015). 39. Ho, Y. S. & McKay, G. Pseudo-second order model for sorption processes. *Process Biochem.* 34, 451–465, [https://doi.org/10.1016/S0032-9592\(98\)00112-5](https://doi.org/10.1016/S0032-9592(98)00112-5) (1999). 40. Elmorsi, R. R. et al. Adsorption of methylene blue and Pb²⁺ by using acid-activated *Posidonia oceanica* waste. *Sci. Rep.* 9, 3356, <https://doi.org/10.1038/s41598-019-39945-1> (2019). 41. Ma, T., Chang, P. R., Zheng, P., Zhao, F. & Ma, X. Fabrication of ultra-light graphene-based gels and their adsorption of methylene blue. *Chem. Eng. J.* 240, 595–600, <https://doi.org/10.1016/j.cej.2013.10.077> (2014). 42. McCoy, T. M. et al. Bulk properties of aqueous graphene oxide and reduced graphene oxide with surfactants and polymers: adsorption and stability. *Phys. Chem. Chem. Phys.* 20, 16801–16816, <https://doi.org/10.1039/C8CP02738B> (2018). 43. Ma, Y.-X. et al. One-step fabrication of β -cyclodextrin modified magnetic graphene oxide nanohybrids for adsorption of Pb(II), Cu(II) and methylene blue in aqueous solutions. *Appl. Surf. Sci.* 459, 544–553, <https://doi.org/10.1016/j.apsusc.2018.08.025> (2018). 44. Zhu, B., Xia, P., Ho, W. & Yu, J. Isoelectric point and adsorption activity of porous g-C₃N₄. *Appl. Surf. Sci.* 344, 188–195, <https://doi.org/10.1016/j.apsusc.2015.03.086> (2015). 45. Liu, H. et al. Fabrication of surface alkalized g-C₃N₄ and TiO₂ composite for the synergistic adsorption-photocatalytic degradation of methylene blue. *Appl. Surf. Sci.* 473, 855–863, <https://doi.org/10.1016/j.apsusc.2018.12.162> (2019). 46. Su, H., Li, W., Han, Y. & Liu, N. Magnetic carboxyl functional nanoporous polymer: synthesis, characterization and its application for methylene blue adsorption. *Sci. Rep.* 8, 6506, <https://doi.org/10.1038/s41598-018-24873-3> (2018). 47. Wei, X. et al. Different adsorption-degradation behavior of methylene blue and Congo red in nanoceria/H₂O₂ system under alkaline conditions. *Sci. Rep.* 9, 4964, <https://doi.org/10.1038/s41598-018-36794-2> (2019). 48. Yang, X. et al. Rapid degradation of methylene blue in a novel heterogeneous Fe₃O₄@rGO@TiO₂-catalyzed photo-Fenton system. *Sci. Rep.* 5, 10632, <https://doi.org/10.1038/srep10632> (2015).

Acknowledgements This work was supported by the Ministry of Research, Technology, and Higher Education (RISTEKDIKTI) Indonesia through Student Creativity Program 2019 (RR, GA) and

23 **Research Grant No.** 200B /WM01.5/N/

2019 (WI). Author contributions R.W. and G.A. carried out the experiments and wrote the main manuscript. W.I. and S.P.S. finalized and revised the

020-57433-9 4 Scientific Reports | (2020) 10:667 |

[2https://doi.org/10.1038/s41598-](https://doi.org/10.1038/s41598-)

020-57433-9 5 Scientific Reports | (2020) 10:667 |

[2https://doi.org/10.1038/s41598-](https://doi.org/10.1038/s41598-)

020-57433-9 Scientific Reports | (2020) 10:667 |

[2https://doi.org/10.1038/s41598-](https://doi.org/10.1038/s41598-)

020-57433-9 7 Scientific Reports | (2020) 10:667 | <https://doi.org/10.1038/s41598-020-57433-9> 8 Scientific Reports | (2020) 10:667 |

[2https://doi.org/10.1038/s41598-](https://doi.org/10.1038/s41598-)

020-57433-9 9